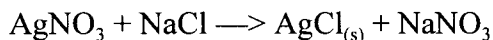


For each of the following equations, write the formulas for the compounds first and then balance the equations. On the line to the right, write the type of reaction, or no reaction for those which do not occur.

1. silver nitrate + sodium chloride $\longrightarrow$ silver chloride + sodium nitrate



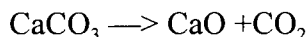
Double replacement

2. zinc + bromine gas $\longrightarrow$ zinc bromide



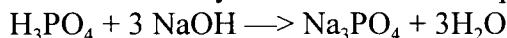
Synthesis

3. calcium carbonate $\longrightarrow$ calcium oxide + carbon dioxide



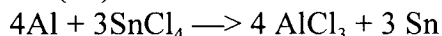
Decomposition

4. phosphoric acid + sodium hydroxide $\longrightarrow$ sodium phosphate + water



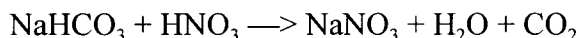
Water-forming

5. aluminum + tin(IV) chloride $\longrightarrow$ aluminum chloride + tin



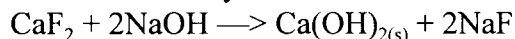
Single Replacement

6. sodium bicarbonate + nitric acid $\longrightarrow$ sodium nitrate + water + carbon dioxide



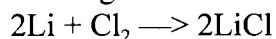
Carbonate

7. calcium fluoride + sodium hydroxide $\longrightarrow$ calcium hydroxide + sodium fluoride



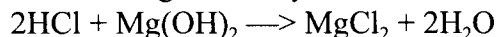
Double Replacement

8. lithium + chlorine gas $\longrightarrow$ lithium chloride



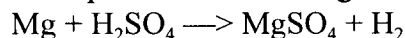
Synthesis

9. hydrochloric acid + magnesium hydroxide $\longrightarrow$ magnesium chloride + water



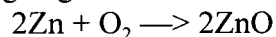
Water-forming

10. magnesium + sulphuric acid $\longrightarrow$ magnesium sulphate + hydrogen gas



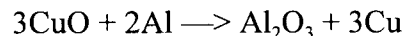
Single replacement

11. zinc + oxygen gas $\longrightarrow$ zinc oxide



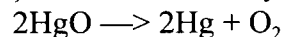
Synthesis (combustion)

12. copper (II) oxide + aluminum $\longrightarrow$ aluminum oxide + copper



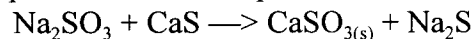
Single replacement

13. mercury (II) oxide $\longrightarrow$ mercury + oxygen gas



Decomposition

14. sodium sulphite + calcium sulphide $\longrightarrow$ sodium sulphide + calcium sulphite



double replacement